

Crystal Data: Hexagonal. *Point Group:* 3. As irregular grains to 5 μm .

Physical Properties: *Cleavage:* None; good parting. *Fracture:* Conchoidal. *Tenacity:* Sectile. Hardness = < 2 D(meas.) = n.d. D(calc.) = 1.364 Easily soluble in water.

Optical Properties: Transparent. *Color:* Colorless. *Streak:* White. *Luster:* Waxy. *Optical Class:* Uniaxial (+). $\omega = 1.402(1)$ $\varepsilon = 1.408(1)$

Cell Data: *Space Group:* R3 (synthetic $\text{Mg}(\text{CH}_3\text{SO}_3)_2 \cdot 12\text{H}_2\text{O}$). $a = 9.27150(8)$ $c = 21.1298(4)$
Z = 3

X-ray Powder Pattern: Calculated pattern.

4.64 (100), 3.87 (89), 3.87 (69), 4.41 (44), 7.04 (42), 6.39 (39), 3.74 (35)

Chemistry:	(1)
Mg	5.64
C	5.58
S	4.89
H	7.03
<u>O</u>	<u>66.86</u>
Total	100.00

(1) $\text{Mg}(\text{CH}_3\text{SO}_3)_2 \cdot 12\text{H}_2\text{O}$; the composition of natural material confirmed by ion chromatography and Raman spectroscopy.

Occurrence: As solid inclusions in an ice core (from a depth of 576.5 m), presumably formed by the fixation of $\text{CH}_3\text{SO}_3\text{H}$ on alkaline particles of marine or continental origin during long-range aerosol transport to the polar region.

Association: Gypsum, ice.

Distribution: Beneath the Dome Fuji station, East Antarctica, near the summit of the eastern Dronning Maud Land plateau.

Name: Honors Ernst A.J. Burke (b. 1943), Vrije Universiteit Amsterdam, Netherlands, for his contributions to the mineralogy of opaque minerals and Raman spectrometry of fluid inclusions.

Type Material: In a cold room (at -50°C), Institute of Low Temperature Science, Hokkaido University, Sapporo, Japan (81,616).

References: (1) Güner, F.E.G., T. Sakurai, and T. Hondoh (2013) Ernstburkeite, $\text{Mg}(\text{CH}_3\text{SO}_3)_2 \cdot 12\text{H}_2\text{O}$, a new mineral from Antarctica. *European Jour. Mineral.*, 25, 79-84. (2) (2014) *Amer. Mineral.*, 99, 1513-1514 (abs. ref. 1). (3) Genceli, F.E., M. Lutz, T. Sakurai, A.L. Spek, and T. Hondoh (2010) Crystallization and characterization of magnesium methanesulfonate hydrate $\text{Mg}(\text{CH}_3\text{SO}_3)_2 \cdot 12\text{H}_2\text{O}$. *Crystal Growth & Design*, 10, 4327-4333.